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You do not have to comply with the license for elements of the material in the public domain or where your use is permitted by an applicable exception or limitation . No warranties are given. The license may not give you all of the permissions necessary for your intended use. For example, other rights such as publicity, privacy, or moral rights may limit how you use the material. Atoms are the basic building blocks of everything in the world.Learning about atoms helps us understand the world and its makeup better. Everything in the world consists of atoms, so it's good to know something about them. Here are 10 interesting and useful atom facts.There are three parts to an atom. Protons have a positive electrical charge and are found together with neutrons (no electrical charge) in the nucleus of each atom. Negatively charged electrons orbit the nucleus.Atoms are the smallest particles that make up elements. Each element contains a different number of protons. For example, all hydrogen atoms have one proton while all carbon atoms have six protons.Some matter consists of one type of atom (e.g., gold), while other matter is made of atoms bonded together to form compounds (e.g., sodium chloride).Atoms are mostly empty space. The nucleus of an atom is extremely dense and contains nearly all of the mass of each atom. Electrons contribute very little mass to the atom (it takes 1,836 electrons to equal the size of a proton)and orbit so far away from the nucleus that each atom is 99.9% empty space. If the atom was the size of a sports arena, the nucleus would be the size of a pea. Although the nucleus is much denser compared with the rest of the atom, it too consists mainly of empty space.There are over 100 different kinds of atoms. About 92 of them occur naturally, while the remainder are made in labs. The first new atom made by man was technetium, which has 43 protons.New atoms can be made by adding more protons to an atomic nucleus. However, these new atoms (elements) are unstable and decay into smaller atoms instantaneously. Usually, we only know a new atom was created by identifying the smaller atoms from this decay.The components of an atom are held together by three forces. Protons and neutrons are held together by the strong and weak nuclear forces. Electrical attraction holds electrons and protons. While electrical repulsion repels protons away from each other, the attracting nuclear force is much stronger than electrical repulsion. The strong force that binds together protons and neutrons is 1,038 times more powerful than gravity, but it acts over a very short range, so particles need to be very close to each other to feel its effect.The word "atom" comes from the Greek word for "uncuttable" or "undivided." The name comes from the 5th century BCE Greek philosopher Democritus, who believed matter consisted of particles that could not be cut into smaller particles. For a long time, people believed atoms were the fundamental "uncuttable" unit of matter. While atoms are the building blocks of elements, they can be divided into still smaller particles. Also, nuclear fission and nuclear decay can break atoms into smaller atoms.Atoms are very small. The average atom is about one-tenth of abillionth of a meter across. The largest atom (cesium) is approximately nine times bigger than the smallest atom (helium).Although atoms are the smallest unit of an element, they consist of even tinier particles called quarks and leptons. An electron is a lepton. Protons and neutrons consist of three quarks each.The most abundant type of atom in the universe is the hydrogen atom. Nearly 74% of the atoms in the Milky Way galaxy are hydrogen atoms.You have around 7 billion billion billion atoms in your body, yet you replace about 98% of them every yeara fascinating fact abut atoms! 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Login to create quiz, word search, matching games, or worksheets. If you are not a registered user register here to login. Atoms are the basic building blocks of everything in the world.Learning about atoms helps us understand the world and its makeup better. Everything in the world consists of atoms, so it's good to know something about them. Here are 10 interesting and useful atom facts.There are three parts to an atom. Protons have a positive electrical charge and are found together with neutrons (no electrical charge) in the nucleus of each atom. Negatively charged electrons orbit the nucleus.Atoms are the smallest particles that make up elements. Each element contains a different number of protons. For example, all hydrogen atoms have one proton while all carbon atoms have six protons.Some matter consists of one type of atom (e.g., gold), while other matter is made of atoms bonded together to form compounds (e.g., sodium chloride).Atoms are mostly empty space. 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The interior of the nucleus contains positively charged protons, and almost all atoms' nuclei (with the exception of hydrogen-1) contain neutrally charged neutrons. Nearly one hundred percent of the mass of an atom (99.94%) is contained in the nucleus. The mass of the protons and the mass of the neutrons are almost the same. Electromagnetic force binds the electron cloud to the nucleus. A nearly identical force can allow atoms to bind together, forming molecules. Atoms exist as either electrically neutral, or ions. In an electrically neutral atom, the number of protons equals the number of electrons, thus cancelling each other out. In an ion, however, there are more of one of particle than the other, making the atom have either a positive or negative charge. Atoms are classified according to their numbers of protons or neutrons. The number of protons in the atom will determine its chemical element, and the number of neutrons will determine its isotope. Every element has at least one isotope, and many elements have multiple isotopes. The isotopes undergo radioactive decay due to their unstable nuclei. The scientific field of quantum mechanics has led to a successful model of the atom and its observable properties. It wasn't until the end of the 19th century and beginning of the 20th century that scientists began to discover subatomic particles contained within the atom. As researchers discovered the existence and properties of subatomic particles, a debate ensued about the fact that there was actually something smaller than an atom and that atoms could be divided, nullifying its name. Some proponents wanted to change the name of the atom to reflect this new understanding, but it didn't catch on. The first understandings of the existence of atoms date back to ancient Greece, and it was Democritus who first used the term "atomos" (atom). Early scientists in ancient India independently developed theories on the existence of atoms. In 1661, Robert Boyle first published the theory that all matter was composed of atoms. The nature and structure of atoms was improved upon beginning in the 18th century with the advances in chemistry. Related Links: Facts Chemistry Facts Animals Facts Chemistry Formulas Atoms Facts Enjoy sharper detail, more accurate color, lifelike lighting, believable backgrounds, and more with our new model update. Your generated images will be more polished thanever.See What's NewExplore how consumers want to see climate stories told today, and what that means for yourvisuals.Download Our Latest VisualGPS ReportData-backed trends. Generative AI demos. Answers to your usage rights questions. Our original video podcast covers it allnow ondemand.Watch NowEnjoy sharper detail, more accurate color, lifelike lighting, believable backgrounds, and more with our new model update. 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